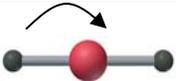
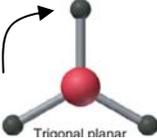
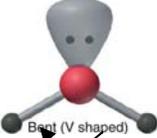
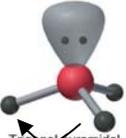
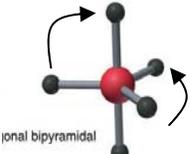
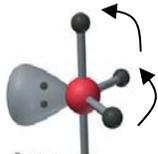
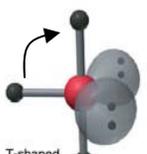
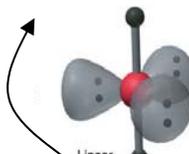
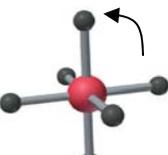
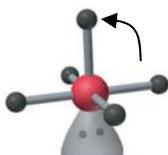
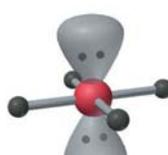


VSEPR Geometries					
Electron Group No. and Geometry	0 lone pairs Molecular Geometry	1 lone pair Molecular Geometry	2 lone pairs Molecular Geometry	3 lone pairs Molecular Geometry	4 lone pairs Molecular Geometry
2 Linear	180° Linear 				
3 Trigonal Planar	120° Trigonal Planar 	< 120° Bent* 	<i>Shapes marked with a * at the end of the name, are ALWAYS polar (even if the atoms surrounding the central atom are identical).</i>		
4 Tetrahedral	109.5° Tetrahedral 	< 109.5° Trigonal Pyramidal* 			
5 Trigonal Bipyramidal	90°, 120° Trigonal bipyramidal 	< 90°, < 120° Seesaw* 	< 90° T-shaped* 	180° Linear 	
6 Octahedral	90° Octahedral 	< 90° Square pyramidal* 	90° Square Planar 	T-shape*	Linear